

(19)



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Office européen des brevets



(11)

EP 1 050 571 A1

(12)

EUROPEAN PATENT APPLICATION

(43) Date of publication:
08.11.2000 Bulletin 2000/45

(51) Int. Cl.⁷: **C10G 45/60, B01J 23/85**

(21) Application number: **00107551.4**

(22) Date of filing: **07.04.2000**

(84) Designated Contracting States:
**AT BE CH CY DE DK ES FI FR GB GR IE IT LI LU
MC NL PT SE**
Designated Extension States:
AL LT LV MK RO SI

(30) Priority: **07.05.1999 IT MI991002**

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(54) **Catalytic composition for the upgrading of hydrocarbons having boiling points within the naphtha range**

(57) A process is described for the upgrading of hydrocarbon mixtures which boil within the naphtha range containing sulfur impurities, i.e. a hydrodesulfuration process with contemporaneous skeleton isomerization and reduced hydrogenation degree of the olefins contained in said hydrocarbon mixtures, the whole process being carried out in a single step. The process is carried out in the presence of a catalytic system comprising a metal of group VI B, a metal of group VIII and a carrier of acid nature consisting of a mesoporous silico-alumina.

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Description

[0001] The present invention relates to a process for the upgrading of hydrocarbon mixtures which boil within the naphtha range containing sulfur impurities, i.e. a hydrodesulfuration process with contemporaneous skeleton isomerization and reduced hydrogenation degree of the olefins contained in said hydrocarbon mixtures, the whole process being carried out in a single step. The process is effected in the presence of a catalytic system comprising a metal of group VI B, a metal of group VIII and a carrier of acid nature consisting of a mesoporous silico-alumina.

[0002] This catalytic system can be used, in particular, for the upgrading of mixtures of hydrocarbons which boil within the naphtha range deriving from cracking processes, preferably mixtures of hydrocarbons having a boiling point within the naphtha range deriving from FCC catalytic cracking (Fluid Catalytic Cracking).

[0003] Hydrocarbons which boil within the naphtha range deriving from FCC (i.e. gasoline cut) are used as blending component of gasolines. For this purpose, it is necessary for them to have a high octane number together with a low sulfur content, to conform with the law restrictions which are becoming increasingly more severe, in order to reduce the emission of contaminants. The sulfur present in gasoline mixtures in fact mainly comes (>90%) from the gasoline cut deriving from FCC.

[0004] This cut is also rich in olefins which have a high octane number. Hydrogenation processes used for desulfuration also hydrogenate the olefins present with a consequent considerable reduction in the octane number (RON and MON). The necessity has therefore been felt for finding a catalytic system which decreases the sulfur content in hydrocarbon mixtures which boil within the naphtha range and, at the same time, minimizes the octane loss (RON and MON), which can be achieved, for example, by the skeleton isomerization of the olefins present and/or by inhibiting the hydrogenation of the olefinic double bond.

[0005] The use of zeolites with a medium pore dimension as isomerization catalysts and the consequent recovery of octane in the charges already subjected to desulfuration are already known (US 5298150, US 5320742, US 5326462, US 5318690, US 5360532, US 5500108, US 5510016, US 5554274, US 599439). In these known processes, in order to obtain hydrodesulfuration with a reduced octane loss, it is necessary to operate in two steps, using in the first step catalysts suitable for desulfuration and in the second step catalysts for recovering the octane number.

[0006] US 5.378.352 describes a process in a single step for desulfurating hydrocarbon fractions, with boiling points within the range of gasolines, using a catalyst which comprises a metal of group VIII, a metal of group VI B and a zeolite.

[0007] MI97A 002288 describes a catalytic system capable of isomerizing n-paraffins having a number of carbon atoms higher than 15 which comprises:

a) a carrier of acid nature consisting of a silica and alumina gel, amorphous to X-rays, with a molar ratio $\text{SiO}_2/\text{Al}_2\text{O}_3$ ranging from 30/1 to 500/1, having a surface area of 500 to 1000 m^2/g , a porosity ranging from 0.3 to 0.6 ml/g and a pore diameter within the range of 10-40 Å;

b) a mixture of metals belonging to groups VI B and VIII deposited on the carrier in an overall quantity of 2 to 50% by weight with respect to the total of (a) + (b).

[0008] The acid carrier of the catalyst preferably has a ratio $\text{SiO}_2/\text{Al}_2\text{O}_3$ ranging from 50/1 to 300/1 and a porosity of 0.4 to 0.5 ml/g .

[0009] The mixture of metals (b) preferably consists of a metal of group VI B selected from molybdenum and tungsten, in a quantity ranging from 5 to 35% by weight, and a non-noble metal of group VIII selected from nickel and cobalt, in a quantity ranging from 0.1 to 5% by weight.

[0010] The carrier based on silica and alumina gel can be conveniently prepared according to what is described in the patent US 5,049,536 or in patent application EP 659478. This carrier can be used as such for the deposition of the metal phase (b) or in extruded form, as described for example in EP 550922 and EP 665055. The metal phase (b) of the catalyst can be introduced by means of aqueous impregnation by wetting the carrier, or extruded carrier, with an aqueous solution of a compound of a metal of group VI B, and drying the resulting product which is then impregnated with an aqueous solution of a compound of the metal of group VIII and dried again. Calcination is then carried out in an oxidating atmosphere at a temperature ranging from 200 to 600°C. Alternatively, a single aqueous solution containing both the compounds of the metals of groups VI B and VIII can be used for contemporaneously introducing these metals. In the case of alcohol impregnation, the acid carrier is suspended in an alcohol solution containing both compounds of the metals of groups VI B and VIII. After impregnation, the solid is dried and calcined.

[0011] These catalytic compositions are useful in the hydroisomerization of n-paraffins having a number of carbon atoms higher than 15 to obtain bases for lubricating oils characterized by a low "pour point" and a high viscosity index.

[0012] It has now been surprisingly found that with the use of these catalytic compositions, optionally also containing a higher percentage of metals, it is possible to desulfurate, with high conversion values, mixtures of hydrocarbons that boil within the naphtha range containing sulfur and olefins and contemporaneously obtain the skeleton isomeriza-

tion of the olefins present with a low hydrogenation degree of the olefinic double bond. This catalytic system is also active at temperatures and pressures that are lower than those preferably used in the known art for hydrodesulfuration.

[0013] Both skeleton isomerization and reduced olefinic hydrogenation enable hydrocarbon mixtures to be obtained, which boil within the naphtha range with very low RON (research octane number) and MON (motor octane number) losses.

[0014] The results obtained do not only relate to the desulfuration of hydrocarbon cuts that boil within the "heavy naphtha" range (130°-250°C), i.e. cuts poor in olefins, but also feeds of "full range naphtha", which boil within the range of 35°-250°C, i.e. in the case of cuts rich in olefins. In fact, the process of the present invention gives a high selectivity for desulfuration with a reduced hydrogenating activity, which represents an additional advantage in terms of octane recovery in the end-gasoline.

[0015] A first object of the present invention therefore relates to a hydrodesulfuration process of hydrocarbon mixtures having boiling ranges within the range of C₄ to 250°C, containing olefins and at least 150 ppm of sulfur, with the contemporaneous skeleton isomerization of these olefins, which comprises putting said hydrocarbon mixtures in contact, in the presence of hydrogen, with a catalytic composition comprising:

a) a carrier of acid nature consisting of a silica and alumina gel, amorphous to X-rays, with a molar ratio SiO₂/Al₂O₃ ranging from 30/1 to 500/1, having a surface area ranging from 500 to 1000 m²/g, a porosity ranging from 0.3 to 0.6 ml/g and a pore diameter within the range of 10-40 Å;

(b) a mixture of metals belonging to groups VI B and VIII deposited on the carrier in an overall quantity of 2 to 67% by weight with respect to the total weight of (a) + (b).

[0016] The acid carrier of the catalyst preferably has a ratio SiO₂/Al₂O₃ ranging from 50/1 to 300/1 and a porosity of 0.4 to 0.5 ml/g.

[0017] According to a preferred aspect of the present invention, the mixture of metals (b) consists of a metal of group VI B, preferably selected from molybdenum and tungsten, and a metal of group VIII, preferably selected from cobalt and nickel.

[0018] According to a particularly preferred aspect of the present invention the mixture of metals (b) consists of molybdenum and cobalt.

[0019] The metal of group VI B is preferably in a quantity ranging from 5 to 50% by weight with respect to the total weight of (a) + (b), even more preferably in a quantity ranging from 8 to 30% by weight. The metal of group VIII is in a quantity ranging from 0.5 to 10% by weight with respect to the total of (a) + (b), even more preferably in a quantity ranging from 1 to 5% by weight.

[0020] The weight percentages of the metal of group VI B and of the metal of group VIII refer to the metal content expressed as metal element of group VI B and metal element of group VIII; in the end-catalyst, after calcination, these metals are in the form of an oxide. According to a particularly preferred aspect the molar ratio between the metal of group VIII and the metal of group VI is less than or equal to 2, preferably less than or equal to 1.

[0021] The silica and alumina gel carrier can be conveniently prepared as described in US 5049536, in EP 659,478 or in EP 812804. In particular, an aqueous solution is prepared, of a tetra-alkyl ammonium hydroxide (TAA-OH), wherein alkyl is for example n-propyl or n-butyl, a soluble compound of aluminum capable of hydrolyzing in Al₂O₃ and a soluble compound of silicon capable of hydrolyzing in SiO₂, the quantity of reagents in solution being such as to respect the following molar ratios:

SiO₂/Al₂O₃ from 30/1 to 500/1;
TAA-OH/SiO₂ from 0.05/1 to 0.2/1;
H₂O/SiO₂ from 5/1 to 40/1.

[0022] The solution thus obtained is heated to gelation, the gel obtained is dried and is calcined in an inert atmosphere and then in an oxidating atmosphere.

[0023] The acid carrier (a) of the catalyst which is used in the process of the present invention can be used as such or in the form of an extruded product with traditional ligands, such as for example aluminum oxide, bohemite or pseudoboehmite. The carrier and ligand can be premixed in weight ratios ranging from 30:70 to 90:10, preferably from 50:50 to 70:30. At the end of the mixing, the product obtained is consolidated into the desired end-form, for example into the form of extruded cylinders or tablets. According to a preferred aspect the silica and alumina gel carrier (a) can be bound with bohemite or pseudoboehmite as described in EP 550,992. Even more preferably the material is used in extruded form as described in EP 665,055, consisting of an inert ligand and silica and alumina gel obtained as follows:

a) preparing an aqueous solution of a tetraalkylammonium hydroxide (TAA-OH), a soluble compound of aluminum capable of hydrolyzing in Al₂O₃ and a silicon compound capable of hydrolyzing in SiO₂, in the following molar

ratios:

SiO₂/Al₂O₃ from 30/1 to 500/1;
TAA-OH/SiO₂ from 0.05/1 to 0.2/1;
H₂O/SiO₂ from 5/1 to 40/1;

b) heating the solution thus obtained to cause hydrolysis and gelation and obtain a mixture A with a viscosity ranging from 0.01 to 100 Pa sec;

c) adding to the mixture A first a ligand belonging to the group of bohemites or pseudoboemites, in a weight ratio with the mixture A of 0.05 to 0.5, and then a mineral or organic acid in a quantity ranging from 0.5 to 8 g per 100 g of ligand;

d) mixing and heating the mixture obtained under point (c) to a temperature ranging from 40° to 90°C until a homogeneous paste is obtained, which is subjected to extrusion;

e) drying of the extruded product and calcination in an oxidating atmosphere.

[0024] As far as the metal phase (b) of the catalyst is concerned, it can be introduced by means of aqueous or alcohol impregnation. According to a first technique, the silica and alumina gel, also in extruded form, is wetted with an aqueous solution of a compound of a metal of group VI B, for example ammonium heptamolibdate, the resulting product is dried, is optionally calcined, and is then impregnated with an aqueous solution of a compound of the metal of group VIII, for example cobalt nitrate. It is then dried and calcined in an oxidating atmosphere at a temperature ranging from 200 to 600°C. Alternatively a single aqueous solution containing both compounds of the metals of groups VI B and VIII can be used for contemporaneously introducing these metals.

[0025] In the case of alcohol impregnation, the acid carrier is suspended in an alcohol solution containing both compounds of the metals of groups VI B and VIII. After impregnation the solid is dried and calcined.

[0026] The catalysts used in the present invention wherein component (b) corresponds to a mixture of metals belonging to groups VI B and VIII, preferably cobalt and molybdenum, deposited on the carrier in an overall quantity of over 50% and less than or equal to 67% by weight with respect to the total of (a) + (b), are new and are a further object of the present invention. According to a preferred aspect these new catalytic compositions contain the metal of group VI B in a quantity greater than 45% and less than or equal to 57% by weight with respect to the total weight of (a) + (b), and the metal of group VIII in a quantity ranging from 5 to 10% by weight with respect to the total weight of (a) + (b).

[0027] A further object of the present invention also relates to a catalytic composition comprising:

a) a carrier of acid nature consisting of a silica and alumina gel, amorphous to X-rays, with a molar ratio SiO₂/Al₂O₃ ranging from 30/1 to 500/1, having a surface area ranging from 500 to 1000 m²/g, a porosity ranging from 0.3 to 0.6 ml/g and a pore diameter within the range of 10-40 Å;

(b) a mixture of cobalt and molybdenum deposited on the carrier in an overall quantity of 2 to 50% by weight with respect to the total weight of the catalytic composition (a) + (b).

[0028] The process of the present invention is preferably carried out at a temperature ranging from 220° to 360°C, even more preferably from 250° to 330°C, at a pressure ranging from 5 to 20 kg/cm², even more preferably from 5 to 10 kg/cm², at a WHSV ranging from 1 to 10 hours⁻¹, even more preferably from 2 to 6 hours⁻¹. The quantity of hydrogen ranges from 100 to 500 times the quantity of hydrocarbons present (N/I), even more preferably from 200 to 400 times.

[0029] The hydrocarbon mixture which can be desulfurated according to the present invention contains more than 150 ppm of sulfur. For example hydrocarbon mixtures with a sulfur content of more than 600 ppm, or even higher than 10000 ppm can be subjected to hydrosulfuration.

The hydrocarbon mixtures which are subjected to hydrosulfuration according to the process of the present invention are mixtures having boiling ranges within the range of C₄ to 250°C, wherein C₄ refers to the boiling temperature of a mixture of hydrocarbons having four carbon atoms, and hydrocarbon mixtures boiling within the naphtha range, i.e. having boiling ranges within the range of C₅ to about 220°C, wherein C₅ refers to the boiling temperature of a mixture of hydrocarbons with five carbon atoms, are preferably subjected to hydrosulfuration.

[0030] The catalysts of the present invention are activated, before use, by sulfidation according to the known methods. For example the sulfidation process can be effected in a reducing atmosphere of H₂S/H₂ at a temperature of 300-500°C or by treatment with carbon sulfide in a reducing atmosphere again at 300-500°C.

Example 1**Preparation of catalyst A**

- 5 [0031] 2 g of aluminum isopropylate are dissolved at room temperature in 68.5 g of aqueous solution of tetrapropylammonium hydroxide (TPA-OH at 13.35% by weight). The solution is heated to 60°C and 104.1 g of tetraethylsilicate (TES) are added. The resulting mixture has the following ratios: $\text{SiO}_2/\text{Al}_2\text{O}_3 = 102$, $\text{TPA-OH}/\text{SiO}_2 = 0.09$
- [0032] On maintaining this mixture under stirring at 60°C for 40 minutes, a homogeneous gel is formed which is dried in a stream of air at 90°C and then calcined at 550°C in a stream of nitrogen for 3 hours and subsequently a
- 10 stream of air for a further 10 hours at the same temperature. A silica and alumina gel carrier (a) is obtained, which is amorphous to X-rays, with a quantitative yield with respect to the materials initially charged, with a pore diameter ranging from 10 to 40 Å, a porosity of 0.47 ml/g.
- [0033] The gel thus obtained is bound with pseudoboehmite, the latter in a quantity of 39% by weight with respect to the total weight of the calcined silica and alumina gel plus the ligand, extruded into cylindrical pellets and ground (40-
- 15 70 mesh, $A_{\text{sup}} = 660 \text{ m}^2/\text{g}$). 10 g of the material thus obtained are then impregnated with 25 ml of aqueous solution containing 10.3 g of $(\text{NH}_4)_6\text{Mo}_7\text{O}_{24} \cdot 4\text{H}_2\text{O}$ (Ammonium heptamolibdate, hereafter called EMA) and left to rest at room temperature for 20 hours. The mixture is then dried in an oven in air at 110°C for 2 hours. The dried product is subsequently impregnated with 12 ml of aqueous solution containing 1.17 g of $\text{Co}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ (Cobalt nitrate, hereafter called CoN), the whole mixture being left to rest at room temperature for 20 hours. It is then dried in an oven in air at
- 20 110°C for 1.5 hours and calcined at 500°C for 4 hours, in air (rising rate: 180°C/hour). The chemical analysis of catalyst A relating to the metal content is indicated in Table 1.

Example 2**Preparation of catalyst A1**

- [0034] 2 g of aluminum isopropylate are dissolved at room temperature in 68.5 g of aqueous solution of tetrapropylammonium hydroxide (TPA-OH at 13.35% by weight). The solution is heated to 60°C and 104.1 g of tetraethylsilicate (TES) are added. The resulting mixture has the following ratios:
- 30 $\text{SiO}_2/\text{Al}_2\text{O}_3 = 102$, $\text{TPA-OH}/\text{SiO}_2 = 0.09$

- [0035] The temperature is maintained at 60-65°C until a mixture A having a viscosity of 0.011 Pa sec, is obtained. This mixture, after aging at room temperature for 20 hours, is put in a mixer and pseudoboehmite Versal 150 (La Roche)
- 35 in a quantity of 39% by weight with respect to the end weight of extruded silica and alumina gel, component a) with ligand, and 2.16 g of methylcellulose (Methocel Fluka 64625), are added, under mixing. After about an hour of mixing 0.26 g of glacial acetic acid are added and the temperature of the paste-mixer jacket is brought to about 50-60°C. The mixing is containing under heating until a homogeneous paste having a consistency suitable for extrusion, is obtained. At the end of the extrusion, aging is carried out for a night at room temperature, followed by drying at 100°C for 5 hours and calcination at 550°C for 8 hours in air. 10 g of the extruded silica and alumina gel, component a) with ligand thus
- 40 obtained are then impregnated with 25 ml of an aqueous solution containing 10.3 g of $(\text{NH}_4)_6\text{Mo}_7\text{O}_{24} \cdot 4\text{H}_2\text{O}$ (Ammonium heptamolibdate, hereafter called EMA) and left to rest at room temperature for 20 hours. The mixture is then dried in an oven in air at 110°C for 2 hours. The dried product is subsequently impregnated with 12 ml of aqueous solution containing 1.17 g of $\text{Co}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ (Cobalt nitrate, hereafter called CoN), the whole mixture being left to rest at
- 45 room temperature for 20 hours. It is then dried in an oven in air at 110°C for 1.5 hours and calcined at 500°C for 4 hours, in air (rising rate: 180°C/hour). The chemical analysis of catalyst A1 relating to the metal content is the same as that for catalyst A.

Example 3**Preparation of catalyst B**

- [0036] 10 g of the silica and alumina gel prepared, extruded and ground as described in example 1 are impregnated with 22 ml of aqueous solution containing 3.5 g of EMA and left to rest at room temperature for 20 hours. The mixture
- 55 is then dried in an oven in air at 100°C, for 2 hours. The dried product is subsequently impregnated with 12 ml of aqueous solution containing 2.46 g of CoN, the whole mixture being left to rest at room temperature for 23 hours. It is then dried in an oven in air at 100°C for 6 hours and calcined at 550°C for 4 hours, in air (rising rate: 180°C/hour). The chemical analysis of the catalyst relating to the metal content is indicated in Table 1.

Example 4**Preparation of catalyst C**

- 5 [0037] 13.3 g of the silica and alumina gel prepared, extruded and ground as described in example 1 are impregnated with 29.2 ml of aqueous solution containing 2.5 g of EMA and left to rest at room temperature for 18 hours. The mixture is then dried in an oven in air at 100°C, for 2 hours. The dried product is subsequently impregnated with 29.2 ml of aqueous solution containing 1.75 g of CoN, the whole mixture being left to rest at room temperature for 20 hours. It is then dried in an oven in air at 100°C for 2 hours and calcined at 550°C for 4 hours, in air (rising rate: 180°C/hour).
- 10 The chemical analysis of the catalyst relating to the metal content is indicated in Table 1.

Example 5**Preparation of catalyst D**

- 15 [0038] The same procedure is adopted as in example 4, but using an acid silica and alumina gel carrier with a ratio $\text{SiO}_2/\text{Al}_2\text{O}_3 = 204$, extruded but not ground. This carrier is prepared as described in example 1 using a quantity of TES equal to 208.2 g.

Example 6**Preparation of catalyst E**

- 20 [0039] The same procedure is adopted as in example 5, except that the silica and alumina gel carrier is ground after extrusion to 40-70 mesh.

Example 7**Comparative catalyst F**

- 30 [0040] A commercial catalyst is used, containing cobalt and molybdenum impregnated on γ -alumina. The chemical analysis of the catalyst relating to the metal content is indicated in Table 1.

Example 7a**Preparation of catalyst G - comparative**

- [0041] A comparative catalyst is prepared by introducing cobalt and molybdenum during the synthesis of the acid silica and alumina gel carrier as described in EP 748652.
- 40 [0042] 3.3 g of CoN are dissolved in 47.48 g of BuOH, under continuous stirring at 60°C for 15 minutes; 1.0 g of $\text{Si}(\text{OC}_2\text{H}_5)_4$ (tetraethylorthosilicate, TES) and 25.1 g of $\text{Al}(\text{OC}_4\text{H}_9)_3$ (aluminum sec-butoxide) are then added, and the stirring continued at 60°C, for 15 minutes. Suspension G1 is obtained.
- [0043] 3.2 g of EMA are dissolved in 33.0 g of $(\text{C}_3\text{H}_7)_4\text{NOH}$ (Tetrapropylammonium hydroxide, solution at 19.2%) at room temperature, obtaining solution G2)
- 45 [0044] Solution G2) is poured into suspension G1), under stirring, at 80°C, these conditions being maintained for 1 h. This is followed by aging at room temperature for 16 h, drying in a vacuum oven at 100°C for 6 h, and calcination in muffle according to the following temperature program: heating to 200°C (5°C/min); pause at 200°C for 2 h; heating to 550°C (5°C/min); pause at 550°C for 3 h; spontaneous cooling to room temperature.
- [0045] Table 1 summarizes the physico-chemical characteristics and chemical analysis for catalysts A-E according to the process at the present invention and for comparative catalysts F and G.
- 50 [0046] In particular the chemical analysis relates to the metal content expressed as weight percentage of Co and Mo and as their molar ratio:

Table 1

Catal.	Co/Mo	Mo(wt %)	Co (wt %)	A_{sup} (m ² /g)
A	0.069	44.2	1.9	310

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Table 1 (continued)

Catal.	Co/Mo	Mo(wt %)	Co (wt %)	A _{sup} (m ² /g)
B	0.37	16.0	3.6	260
C	0.35	10.7	2.3	370
D	0.45	8.2	2.2	370
E	0.45	8.3	2.3	440
F	0.43	12.0	3.2	245
G	0.61	18.1	6.8	430

Tests on model charge

[0047] The performances of catalysts B, C and E are provided hereunder, obtained using the so-called "model charge" as feed, which represents the typical composition of FCC gasoline, in terms of S content and percentage of olefins. The model charge consists of:

- 30 wt % of 1-pentene;
- 0.25 wt % of thiophene, equal to about 1000 ppm S;
- the complement to 100 is n-hexane.

[0048] The operating conditions are selected so as to obtain conversions higher than 90% (as HDS) and are kept relatively bland (250°C < T < 300°C); H₂ is fed at P_{H2}= 10 bars.

[0049] The parameters selected for evaluating the catalytic activity are:

- HDS conversion % equal to
100 x (ppm S_{in}-ppm S_{out})/S_{in};
- ISO isomerizing activity %, equal to
100 x (i-pentanes + i-pentenenes)/SUM C₅;
- HYD hydrogenating activity %, equal to
100 x n-pentane_{out}/1 pentene_{in};
- HDS/HYD ratio, measurement of the selectivity of the catalyst towards desulfuration;
- HYD/ISO ratio, inversely proportional to the selectivity towards branched products.

[0050] The catalysts are all activated with the same procedure in a stream of H₂S/H₂.

Example 8

Activity of catalyst B

[0051] 2 g of catalyst B prepared according to example 3, diluted with corindone, are charged into a reactor (40-70 mesh) and activated by a stream of the mixture H₂S/H₂ (10% vol H₂S), at 400°C, for 3 hours; the system is then brought under a pressure of H₂ up to 10 bars, with T = 250°C. The model charge is sent to the catalytic bed, with a ratio H₂/hydrocarbon charge of 300N/l. The operating conditions and catalytic results are indicated in Table 2.

Table 2

T(°C)	WHSV (h ⁻¹)	HDS%	ISO%	HYD%	HDS/HYD	HYD/ISO
286	4.06	96.7	37.4	38.3	2.5	1.0
288	6.23	94.4	26.4	38.3	2.5	1.4
269	4.06	89.9	21.5	34.1	2.6	1.6

Example 9**Activity of catalyst C**

5 [0052] 2 g of catalyst C prepared according to example 4, diluted with corindone, are charged into a reactor (40-70 mesh) and activated by a stream of the mixture H_2S/H_2 (10% vol H_2S), at 400°C, for 3 hours; the system is then brought under a pressure of H_2 up to 10 bars, with $T = 250^\circ C$. The model charge is sent to the catalytic bed, with a ratio H_2 /hydrocarbon charge of 300N/l. The operating conditions and catalytic results are indicated in Table 3.

Table 3

T(°C)	WHSV (h^{-1})	HDS%	ISO%	HYD%	HDS/HYD	HYD/ISO
287	6.12	96.3	43.1	30.5	3.2	0.7
271	4.21	94.1	36.1	30.2	3.1	0.8
287	4.21	96.6	34.6	43.1	2.2	1.2

Example 10**Activity of catalyst E**

20 [0053] 2 g of catalyst E prepared according to example 6, diluted with corindone, are charged into a reactor (40-70 mesh) and activated by a stream of the mixture H_2S/H_2 (10% vol H_2S), at 400°C, for 3 hours; the system is then brought under a pressure of H_2 up to 10 bars, with $T = 250^\circ C$. The model charge is sent to the catalytic bed, with a ratio H_2 /hydrocarbon charge of 300N/l. The operating conditions and catalytic results are indicated in Table 4.

Table 4

T(°C)	WHSV (h^{-1})	HDS%	ISO%	HYD%	HDS/HYD	HYD/ISO
287	4.1	98.2	30.1	50.9	1.9	1.7
288	6.1	96.8	26.3	45.3	2.1	1.7
270	4.1	93.3	17.8	44.7	2.1	2.5

Example 11**Activity of catalyst F - comparative**

40 [0054] 2 g of catalyst F prepared according to example 7, diluted with corindone, are charged into a reactor (40-70 mesh) and activated by a stream of the mixture H_2S/H_2 (10% vol H_2S), at 400°C, for 3 hours; the system is then brought under a pressure of H_2 up to 10 bars, with $T = 250^\circ C$. The model charge is sent to the catalytic bed, with a ratio H_2 /hydrocarbon charge of 300N/l. The operating conditions and catalytic results are indicated in Table 5.

Table 5

T(°C)	WHSV (h^{-1})	HDS%	ISO%	HYD%	HDS/HYD	HYD/ISO
250	4.0	89.7	2.5	37.4	2.4	15.0
270	4.0	100.0	2.6	80.6	1.2	31.0

Example 12**Activity of catalyst G - comparative**

55 [0055] 2 g of catalyst F prepared according to example 7a, diluted with corindone, are charged into a reactor (40-

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70 mesh) and activated by a stream of the mixture H_2S/H_2 (10% vol H_2S), at 400°C, for 3 hours; the system is then brought under a pressure of H_2 up to 10 bars, with $T = 250^\circ C$. The model charge is sent to the catalytic bed, with a ratio H_2 /hydrocarbon charge of 300NVI. The operating conditions and catalytic results are indicated in Table 6.

Table 6

T(°C)	WHSV (h^{-1})	HDS%	ISO%	HYD%	HDS/HYD	HYD/ISO
293	3.9	91.5	0.8	86.0	1.1	107.5
274	3.9	92.1	0.7	89.9	1.0	128.4

[0056] These tables demonstrate the greater isomerizing activity of the catalysts of the present invention (B, C and E), both in absolute terms (ISO%) and relating to the hydrogenating capacity (HYD/ISO).

[0057] A similar comparison between the formulations B, C and E, on the one hand, and F and G on the other, can be based on the respective performances under the same operating conditions ($WHSV = 4 \text{ hours}^{-1}$; $T = 270^\circ C$, as indicated in Table 7).

Table 7

Catalyst	HDS (%)	ISO (%)	HDS/HYD	HYD/ISO
B	89.9	21.0	2.6	1.6
C	94.1	36.1	3.1	0.8
E	93.3	17.8	2.1	2.5
F	100.0	2.6	1.2	31.0
G	92.1	0.7	1.0	128.4

[0058] From this comparison it can be seen that, with the same operating conditions, the catalysts of the present invention (B, C and E) have a higher isomerizing activity (more than an order of magnitude) and also a lower hydrogenating activity (lower HYD/ISO ratio), with respect to the comparative catalysts.

[0059] In particular the best performances are obtained with catalyst C having a composition corresponding to a ratio $SiO_2/Al_2O_3 = 100$ and a molybdenum content of 10.6%. This is confirmed by the data at $T = 288^\circ C$ and $WHSV = 6.1 \text{ hours}^{-1}$ (Table 8).

Table 8

Catalyst	HDS (%)	ISO (%)	HDS/HYD	HYD/ISO
B	94.4	26.4	2.5	1.4
C	96.3	43.1	3.2	0.7
E	96.8	26.3	2.1	1.7

Tests on FCC Full Range Naphtha

[0060] Some examples are provided hereunder of performances of catalysts of the present invention, evaluated on FCC Full Range gasoline, characterized by the following composition:

Table 9

S (ppm)	Paraff. (n+i-)	Naphthenes	olefins (n+i-)	Cycloolefins	aromatics	B.P.> 200°C	RON	MON
1360	20.4	9.4	30.3	5.6	24.8	8.9	93.3	80.7

wherein S ppm is the sulfur content and columns two to six indicate the volume percentage of normal and iso paraffins,

naphthenes, normal and iso olefins, cyclo-olefins, aromatics, respectively. The last column indicates the volume percentage of the fraction which boil at over 200°C.

Example 13

Activity of catalyst B

[0061] 2 g of catalyst B prepared according to example 3, diluted with corindone, are charged into a reactor (40-70 mesh) and activated in the presence of a mixture of H_2S/H_2 (10% vol H_2), up to 400°C, for 3 hours; the system is then brought under a pressure of H_2 up to 10 bars, sending the feed consisting of FCC Full Range gasoline having the composition of Table 9, onto the catalyst with a ratio H_2 /hydrocarbon charge of 300N/l and WHSV of 4 hours⁻¹. The operating conditions and catalytic results are indicated in Table 10.

Table 10

S (ppm)	T(°C)	Paraff. (n+i-)	Naphthenes	Olefins (n+i-)	Aromatics	B.P.> 200°C	RON	MON
501	272	26.3	9.8	24.8	25.1	9.4	92.4	81.2
277	288	29.2	10.5	21.6	24.8	9.5	91.0	80.3
127	304	33.3	11.2	18.7	24.4	9.5	90.1	80.2

Example 14

Activity of catalyst C

[0062] 2 g of catalyst C prepared according to example 4, diluted with corindone, are charged into a reactor (40-70 mesh) and activated in the presence of a mixture of H_2S/H_2 (10% vol H_2), up to 400°C, for 3 hours; the system is then brought under a pressure of H_2 up to 10 bars, sending the feed consisting of FCC Full Range gasoline (see Table 9), onto the catalyst with a ratio H_2 /hydrocarbon charge of 300N/l and WHSV = 4.4 h⁻¹.

[0063] The operating conditions and catalytic results are indicated in Table 11.

Table 11

S (ppm)	T(°C)	Paraff. (n+i-)	Naphthenes	Olefins (n+i-)	Aromatics	B.P.> 200°C	RON	MON
538	252	24.6	9.6	26.8	24.8	9.5	---	81.0
224	272	26.2	9.9	24.8	24.9	9.7	91.7	80.9
115	287	27.8	10.4	22.8	25.0	10.0	91.5	80.7

Example 15

Activity of catalyst F - comparative

[0064] 2 g of catalyst F prepared according to example 7, diluted with corindone, are charged into a reactor (40-70 mesh) and activated in the presence of a mixture of H_2S/H_2 (10% vol H_2), up to 400°C, for 3 hours; the system is then brought under a pressure of H_2 up to 10 bars, sending the feed consisting of FCC Full Range gasoline (see Table 9), onto the catalyst, with a ratio H_2 /hydrocarbon charge of 300N/l and WHSV = 4 h⁻¹.

[0065] The operating conditions and catalytic results of the gasoline thus obtained are indicated in Table 12.

Table 12

S (ppm)	T(°C)	Paraff. (n+i-)	Naphthenes	Olefins (n+i-)	Aromatics	B.P.> 200°C	RON	MON
886	230	22.8	10.0	27.8	23.3	10.4	92.3	80.3
551	248	24.2	10.3	26.5	23.2	10.2	91.8	80.2

Table 12 (continued)

S (ppm)	T(°C)	Paraff. (n+i-)	Naphthenes	Olefins (n+i-)	Aromatics	B.P.> 200°C	RON	MON
158	269	28.6	11.3	21.7	23.4	9.7	89.3	79.3

[0066] Table 13 compares the variations in the octane qualities of the products obtained with catalysts B, C and F, at 90% of HDS conversion. The catalysts of the present invention (B and C) have a lower loss, in terms of both RON and MON, with respect to F; catalyst C is, in particular, the one which ensures the lowest loss.

Table 13

Catalyst	Δ RON	Δ MON	Δ RON+MON/2)
B	-3.2	-0.5	-1.85
C	-1.8	0	-0.9
F	-4.0	-1.4	-2.7

Claims

- The hydrodesulfuration of hydrocarbon mixtures having boiling ranges within the range of C_4 to 250°C, containing olefins and at least 150 ppm of sulfur, with the contemporaneous skeleton isomerization of said olefins, which comprises putting these hydrocarbon mixtures in contact with hydrogen and with a catalytic composition comprising:
 - a carrier of an acid nature consisting of a silica and alumina gel, amorphous to X-rays, with a molar ratio SiO_2/Al_2O_3 of 30/1 to 500/1, having a surface area ranging from 500 to 1000 m^2/g , a porosity of 0.3 to 0.6 ml/g and a pore diameter within the range of 10-40 Å;
 - a mixture of metals belonging to groups VI B and VIII deposited on the carrier in an overall quantity ranging from 2 to 67% by weight with respect to the total of (a) + (b).
- The process according to claim 1, wherein the acid carrier of the catalyst has a ratio SiO_2/Al_2O_3 ranging from 50/1 to 300/1 and a porosity of 0.4 to 0.5 ml/g.
- The process according to claim 1, wherein the mixture of metals (b) consists of a metal of group VI B and a metal of group VIII.
- The process according to claim 1 or 3, wherein the metal of group VI B is selected from molybdenum and tungsten, and the metal of group VIII is selected from cobalt and nickel.
- The process according to claim 4, wherein the metal of group VI B is molybdenum and the metal of group VIII is cobalt.
- The process according to claim 1 or 3, wherein the metal of group VI B is in a quantity ranging from 5 to 50% by weight with respect to the total of (a) + (b) and the metal of group VIII is in a quantity ranging from 0.5 to 10% by weight with respect to the total of (a) + (b).
- The process according to claim 6, wherein the metal of group VI B is in a quantity ranging from 8 to 30% by weight and the metal of group VIII is in a quantity ranging from 1 to 5% by weight.
- The process according to claim 1 or 3, wherein the molar ratio between the metal of group VIII and the metal of group VI B is less than or equal to 2.
- The process according to claim 8, wherein the molar ratio is less than or equal to 1.
- The process according to claim 1 or 3, wherein the silica and alumina gel carrier (a) is used in the form of an extruded product with a ligand.
- The process according to claim 10, wherein the ligand is selected from aluminum oxide, bohemite and pseudobo-

hemite.

12. The process according to claim 10, wherein the silica and alumina gel carrier (a) and the ligand are premixed in weight ratios ranging from 30:70 and 90:10 and consolidated into the desired end-form.

13. The process according to claim 10, wherein the silica and alumina gel in extruded form is prepared as follows:

a) preparing an aqueous solution of a tetraalkylammonium hydroxide (TAA-OH), a soluble compound of aluminum capable of hydrolyzing in Al_2O_3 and a silicon compound capable of hydrolyzing in SiO_2 , in the following molar ratios

$\text{SiO}_2/\text{Al}_2\text{O}_3$ from 30/1 to 500/1;
TAA-OH/ SiO_2 from 0.05/1 to 0.2/1;
 $\text{H}_2\text{O}/\text{SiO}_2$ from 5/1 to 40/1;

b) heating the solution thus obtained to cause hydrolysis and gelation and obtain a mixture A with a viscosity ranging from 0.01 to 100 Pa sec;

c) adding to the mixture A first a ligand belonging to the group of bohemites or pseudobohemites, in a weight ratio with the mixture A of 0.05 to 0.5, and then a mineral or organic acid in a quantity ranging from 0.5 to 8 g per 100 g of ligand;

d) mixing and heating the mixture obtained under point (c) to a temperature ranging from 40° to 90°C until a homogeneous paste is obtained, which is subjected to extrusion;

e) drying of the extruded product and calcination in an oxidating atmosphere.

14. The process according to claim 1, carried out at a temperature ranging from 220°C to 360°C, at a pressure ranging from 5 to 20 kg/cm^2 , at a WHSV ranging from 1 to 10 hours^{-1} and with a quantity of hydrogen ranging from 100 to 500 times the quantity of hydrocarbons present (N/I).

15. The process according to claim 14, carried out at a temperature ranging from 250°C to 330°C, at a pressure ranging from 5 to 10 kg/cm^2 , at a WHSV ranging from 2 to 6 hours^{-1} and with a quantity of hydrogen ranging from 200 to 400 times the quantity of hydrocarbons present (N/I).

16. The process according to claim 1, wherein the hydrocarbon mixture which is subjected to desulfuration contains more than 600 ppm of sulfur.

17. The process according to claim 1, wherein the hydrocarbon mixtures which are subjected to hydrodesulfuration are mixtures having boiling ranges within the range of C_5 to 220°C.

18. The process according to claim 1, wherein the catalysts are activated by sulfidation.

19. A bifunctional catalyst comprising:

(a) a carrier of an acid nature consisting of a silica and alumina gel, amorphous to X-rays, with a molar ratio $\text{SiO}_2/\text{Al}_2\text{O}_3$ of 30/1 to 500/1, having a surface area ranging from 500 to 1000 m^2/g , a porosity of 0.3 to 0.6 ml/g and a pore diameter within the range of 10-40 Å;

b) a mixture of metals belonging to groups VI B and VIII deposited on the carrier in an overall quantity which is greater than 50% and less than or equal to 67% by weight with respect to the total of (a) + (b).

20. The catalyst according to claim 19, containing a metal of group VI B in a quantity greater than 45% and less than or equal to 57% by weight with respect to the total of (a) + (b), and a metal of group VIII in a quantity ranging from 5 to 10% by weight with respect to the total of (a) + (b).

21. The catalyst according to claims 19 or 20, wherein the metal of group VI B is Mo and the metal of group VIII is Co.

22. A bifunctional catalyst containing:

(a) a carrier of an acid nature consisting of a silica and alumina gel, amorphous to X-rays, with a molar ratio $\text{SiO}_2/\text{Al}_2\text{O}_3$ of 30/1 to 500/1, having a surface area ranging from 500 to 1000 m^2/g , a porosity of 0.3 to 0.6 ml/g

and a pore diameter within the range of 10-40 Å;

b) a mixture of cobalt and molybdenum deposited on the carrier in an overall quantity ranging from 2 to 50% by weight with respect to the total of (a) + (b).

5 23. A process for preparing the catalysts according to claim 19, which comprises:

i) wetting the silica and alumina gel, optionally in extruded form, with an aqueous solution of a compound of the metal of group VI B;

ii) drying the resulting product, and optionally calcining it;

10 iii) impregnating the product obtained under point ii) with an aqueous solution of a compound of the metal of group VIII;

iv) drying the impregnated product obtained under point iii) and calcining it in an oxidating atmosphere at a temperature ranging from 200 to 600°C.

15 24. A process for preparing the catalysts according to claim 19, which comprises wetting the silica and alumina gel, optionally in extruded form, with an aqueous solution of a compound of a metal of group VI B and a compound of a metal of group VIII, drying the product thus obtained and calcining it in an oxidating atmosphere at a temperature ranging from 200 to 600°C.

20 25. A process for preparing the catalysts according to claim 22, which comprises:

i) wetting the silica and alumina gel, optionally in extruded form, with an aqueous solution of a compound of molybdenum;

ii) drying the resulting product, and optionally calcining it;

25 iii) impregnating the product obtained under point ii) with an aqueous solution of a compound of cobalt;

iv) drying the impregnated product obtained under point iii) and calcining it in an oxidating atmosphere at a temperature ranging from 200 to 600°C.

30 26. A process for preparing the catalysts according to claim 22, which comprises wetting the silica and alumina gel, optionally in extruded form, with an aqueous solution of a compound of molybdenum and a compound of cobalt, drying the product thus obtained and calcining it in an oxidating atmosphere at a temperature ranging from 200 to 600°C.

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EUROPEAN SEARCH REPORT

Application Number
EP 00 10 7551

DOCUMENTS CONSIDERED TO BE RELEVANT			
Category	Citation of document with indication, where appropriate, of relevant passages	Relevant to claim	CLASSIFICATION OF THE APPLICATION (Int.Cl.7)
X	EP 0 908 231 A (AGIP PETROLI) 14 April 1999 (1999-04-14) * the whole document *	22-26	C10G45/60 B01J23/85
Y	—	1-18	
Y	US 3 156 640 A (MARIUS 'T HART) 10 November 1964 (1964-11-10) * the whole document *	1-18	
A	GB 878 035 A (SHELL) 20 September 1960 (1960-09-20) * the whole document *	1-18	
A	GB 2 051 120 A (INST FRANCAIS DU PETROL) 14 January 1981 (1981-01-14) * the whole document *	1-18	
A	EP 0 543 157 A (IDEMITSU KOSAN CO ;PETROLEUM ENERGY CENTER FOUND (JP)) 26 May 1993 (1993-05-26) * the whole document *	19-21	
D,A	EP 0 550 922 A (EURON SPA ;SNAM PROGETTI (IT); ENIRICERCH SPA (IT)) 14 July 1993 (1993-07-14) * the whole document *	1-26	
			TECHNICAL FIELDS SEARCHED (Int.Cl.7)
			C10G B01J
The present search report has been drawn up for all claims			
Place of search THE HAGUE		Date of completion of the search 24 May 2000	Examiner Michiels, P
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EPO FORM 1508 03.02 (P/4001)

**ANNEX TO THE EUROPEAN SEARCH REPORT
ON EUROPEAN PATENT APPLICATION NO.**

EP 00 10 7551

This annex lists the patent family members relating to the patent documents cited in the above-mentioned European search report. The members are as contained in the European Patent Office EDP file on
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24-05-2000

Patent document cited in search report	Publication date	Patent family member(s)	Publication date
EP 0908231 A	14-04-1999	IT 1295300 B	04-05-1999
		AU 8838698 A	29-04-1999
		CA 2246729 A	09-04-1999
		JP 11193249 A	21-07-1999
US 3156640 A	10-11-1964	DE 1443121 A	31-10-1968
		GB 917091 A	
GB 878035 A		DE 1443074 A	31-10-1968
		US 3149180 A	15-09-1964
		NL 105779 C	
		US 3161697 A	15-12-1964
		DE 1248845 B	
		FR 1289850 A	24-08-1962
		GB 935718 A	
		NL 126095 C	
GB 2051120 A	14-01-1981	NL 251684 A	
		FR 2457317 A	19-12-1980
		BE 883332 A	17-11-1980
		DE 3019216 A	04-12-1980
		IT 1140968 B	10-10-1986
		JP 55155090 A	03-12-1980
		NL 8002881 A	25-11-1980
		US 4334982 A	15-06-1982
EP 0543157 A	26-05-1993	JP 5115781 A	14-05-1993
		US 5336394 A	09-08-1994
		US 5280004 A	18-01-1994
EP 0550922 A	14-07-1993	IT 1252647 B	20-06-1995
		DE 69210002 D	23-05-1996
		DE 69210002 T	02-10-1996
		DK 550922 T	03-06-1996
		NO 301967 B	05-01-1998
		US 5342814 A	30-08-1994

EPO FORM P468

For more details about this annex : see Official Journal of the European Patent Office, No. 12/82